

Code: <b>QG102</b>																		
Name: <b>Química Experimental I</b>																		
Name in English: <b>Experimental Chemistry I</b>																		
Name in Spanish: <b>Química Experimental I</b>																		
Subject type: <b>Weekly</b>																		
Approval Type: <b>Grade and frequency</b>																		
Characteristic: <b>Regular</b>																		
Frequency: <b>75%</b>																		
Period Type / Offering period: <b>Semestral/all the periods</b>																		
Requires Final Exam: <b>yes</b>																		
<b>Vectors</b>																		
<table border="1" style="width: 100%; border-collapse: collapse;"> <thead> <tr> <th>T</th><th>L</th><th>P</th><th>O</th><th>PE</th><th>OE</th><th>SL</th><th>WEEKS</th><th>CREDITS</th></tr> </thead> <tbody> <tr> <td><b>1</b></td><td><b>3</b></td><td>-</td><td>-</td><td>-</td><td>-</td><td><b>4</b></td><td><b>15</b></td><td><b>4</b></td></tr> </tbody> </table>	T	L	P	O	PE	OE	SL	WEEKS	CREDITS	<b>1</b>	<b>3</b>	-	-	-	-	<b>4</b>	<b>15</b>	<b>4</b>
T	L	P	O	PE	OE	SL	WEEKS	CREDITS										
<b>1</b>	<b>3</b>	-	-	-	-	<b>4</b>	<b>15</b>	<b>4</b>										
Occurrence on curriculum: <b>04, 08, 12</b>																		
Pre requirement: None																		
<b>Summary:</b> Experiments illustrating the scientific method, concepts of mol and chemical bond, oxidation-reduction, chemical equilibrium, pH, solubility product, preparation, and purification of substances.																		
<b>Program:</b> Introduction. Laboratory Safety. - Physical and chemical phenomena. - Determination of the metal equivalent. - Quantitative study of the reaction of a metal with acid. - Continuous variations method. - Determination of the molecular mass of a volatile liquid by measuring density. - Melting point of a pure substance. - Illustrative reactions of chemical equilibrium. - Determination of the solubility product of silver acetate. - pH measurements. - Chemical equilibrium. - Oxidation-reduction potential. - Titration.																		
<b>Basic Bibliography</b>																		
1) KOTZ, J. C.; TREICHEL JR., P. <b>Química e Reações Químicas</b> , vol. 1 e 2, 4 <sup>a</sup> ed., LTC, Rio de Janeiro, 2002.																		
2) ATKINS, P. W.; JONES, L.; LAVERMAN, L. <b>Princípios de Química</b> , 7 <sup>a</sup> ed., Bookman, Porto Alegre, 2018.																		
3) FLOWERS, P.; ROBINSON, W. R.; Langley, R.; THEOPOLD, K. <b>Chemistry</b> , OpenStax, Houston, 2015 (e-book disponível em: <a href="https://openstax.org/books/chemistry/pages/1-introduction">https://openstax.org/books/chemistry/pages/1-introduction</a> ).																		
<b>Supplementary Bibliography</b>																		
1) BACCAN, N.; GODINHO, O. E. S.; ALEIXO, L. M.; STEIN, E. <b>Introdução à Semimicroanálise Qualitativa</b> , Editora da UNICAMP, Campinas, 1990.																		
2) VOGEL, A. I. <b>Química Analítica Qualitativa</b> , Editora Mestre Jou, São Paulo, 1981.																		
3) VOGEL, A. I. <b>Análise Química Quantitativa</b> , 6 <sup>a</sup> ed., Editora LTC, Rio de Janeiro, 2002.																		
4) PAVIA, D. L.; LAPMAN, G. M.; KRIZ, G. S.; ENGEL, R. G. <b>Introduction to Organic Laboratory Techniques: a Microscale Approach</b> , 4th ed., Thomson Brooks/Cole, Belmont, 2007.																		
5) CHEMKEYS, disponível em <a href="http://www.chemkeys.com">www.chemkeys.com</a> , e-ISSN 2595-7430.																		